

CLASS X GUESS PAPER SCIENCE

METALS AND NON - METALS

1. What is metallic luster?
2. Define malleability. Name the two most malleable metals.
3. Define ductility. Name the two most ductile metals.
4. Why copper and aluminium are used to make cooking vessel?
5. Name the best and poorest conductor of electricity among metals
6. Why are electric wires are coated with rubber or pvc?
7. Why are school bells made of vessels?
8. Name two metals which have low melting point.
9. What are allotropes? Give two examples.
10. What do you observe when copper is heated in air? Write the equation also.
11. Name two metal oxides which are soluble in water. Write equation also
12. Why are sodium and potassium stored under kerosene?
13. What happens when Mg, Al, Zn, Pb are exposed to air at room temperature?
14. Explain the reaction of Fe and Cu with air, along with equations.
15. Why do Na and K burn when put in cold water?
16. Why do Mg and Ca start floating but not burn on putting in cold water?
17. Under what condition do Al, Fe and Zn react with water? Write the equations also.
18. Why copper, silver and gold donot react with water to liberate hydrogen gas?
19. Why hydrogen is not evolved when metals react with dilute nitric acid?
20. Name two metals which liberate hydrogen gas when they react with dilute nitric acid.
21. What do you mean by the reactivity series of metals?
22. Why do metals lose and non metals gain electrons during chemical reaction?
23. Why are noble gases unreactive?
24. Show the formation of sodium chloride by electron dot representation.
25. Why sodium chloride donot conduct electricity in solid state, but does in molten state?
26. How are ionic compounds formed? Explain with examples.
27. Why are ionic compounds have a high melting and boiling point?
28. Why are ionic compounds solids and somewhat hard?
29. Name two compound which are found in aqueous state on the earth naturally?

30. What are minerals and ores?
31. Why do some metals occur in free state in earth's crust?
32. Name two metals which occur in both free and combined states in the earth's crust.
33. Why the ores of many metals are oxides?
34. What makes it possible to reduce the oxides of Zn, Fe, Pb and Cu using carbon?
35. Why we can't reduce aluminium oxide using carbon?
36. Explain how you will obtain mercury from its ore cinnabar?
37. What is gangue?
38. Explain how you will obtain Cu metal from its ore?
39. How will you convert (a) carbonate ore (b) sulphide ore into its corresponding oxides?
40. What do you mean by thermite reaction? Give example and application.
41. Why is iron not used in pure form?
42. Explain electrolytic refining with example.

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